

**Chapter 6 Reaction Practice Group Quiz Answers**

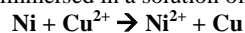
- Magnesium ribbon is burned in oxygen.  
$$2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$$
- A bar of strontium metal is immersed in a 1.0 M copper(II) nitrate solution.  
$$\text{Sr} + \text{Cu}^{2+} \rightarrow \text{Sr}^{2+} + \text{Cu}$$
- A sample of nickel(II) sulfite hexahydrate is heated strongly.  
$$\text{NiSO}_3 \cdot 6\text{H}_2\text{O} \rightarrow \text{NiSO}_3 + 6\text{H}_2\text{O}$$
- Solutions of sodium chromate and lead(II) nitrate are mixed.  
$$\text{CrO}_4^{2-} + \text{Pb}^{2+} \rightarrow \text{PbCrO}_4$$
- Liquid bromine is carefully added to a solution of potassium iodide.  
$$\text{Br}_2 + 2\text{I}^- \rightarrow 2\text{Br}^- + \text{I}_2$$
- A solution of sodium iodide is added to a solution of lead(II) acetate.  
$$2\text{I}^- + \text{Pb}^{2+} \rightarrow \text{PbI}_2$$
- Pure solid phosphorus (P<sub>4</sub>) is burned in air.  
$$\text{P}_4 + 5\text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$$
- Solid sodium carbonate is strongly heated.  
$$\text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{O} + \text{CO}_2$$
- A piece of copper wire is placed in a solution of silver nitrate.  
$$\text{Cu} + \text{Ag}^+ \rightarrow \text{Cu}^+ + \text{Ag}$$
- Solutions of strontium nitrate and sodium sulfate are mixed.  
$$\text{Sr}^{2+} + \text{SO}_4^{2-} \rightarrow \text{SrSO}_4$$
- A small piece of calcium metal is added to hot distilled water.  
$$\text{Ca} + 2\text{HOH} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$$
- Butanol(C<sub>4</sub>H<sub>9</sub>OH) is burned in air.  
$$\text{C}_4\text{H}_9\text{OH} + 6\text{O}_2 \rightarrow 4\text{CO}_2 + 5\text{H}_2\text{O}$$
- A solution of copper(II) chloride is added to a solution of sodium sulfide.  
$$\text{Cu}^{2+} + \text{S}^{2-} \rightarrow \text{CuS}$$
- Solid potassium chlorate is heated strongly.  
$$2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$$
- Powdered strontium oxide is added to distilled water.  
$$\text{SrO} + \text{H}_2\text{O} \rightarrow \text{Sr(OH)}_2$$
- Solutions of cobalt(II) nitrate and sodium hydroxide are mixed.  
$$\text{Co}^{2+} + 2\text{OH}^- \rightarrow \text{Co(OH)}_2$$
- Ethene(C<sub>2</sub>H<sub>4</sub>) gas is burned in air.  
$$\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$$
- A strip of zinc is added to a solution of hydrochloric acid.  
$$\text{Zn} + 2\text{H}^+ \rightarrow \text{Zn}^{2+} + \text{H}_2$$
- Solid nickel(II) sulfide is strongly heated in air.  
$$2\text{NiS} + 3\text{O}_2 \rightarrow 2\text{NiO} + 2\text{SO}_2$$
- Solid pieces of potassium and iodine are heated strongly  
$$2\text{K} + \text{I}_2 \rightarrow 2\text{KI}$$

Name: \_\_\_\_\_

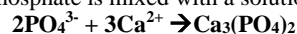
**Honors Chemistry Chapter 6 Reaction Practice Homework Answers**

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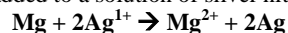
1. A piece of nickel metal is immersed in a solution of copper(II) sulfate.



2. A solution of potassium phosphate is mixed with a solution of calcium acetate.



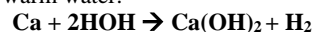
3. A strip of magnesium is added to a solution of silver nitrate.



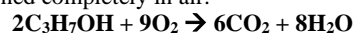
4. Ethyne( $\text{C}_2\text{H}_2$ ) gas is burned in air.



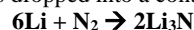
5. Solid calcium is added to warm water.



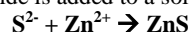
6. Propanol( $\text{C}_3\text{H}_7\text{OH}$ ) is burned completely in air.



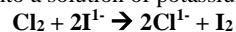
7. A piece of lithium metal is dropped into a container of nitrogen gas.



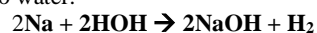
8. A solution of sodium sulfide is added to a solution of zinc nitrate.



9. Chlorine gas is bubbled into a solution of potassium iodide.



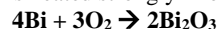
10. Sodium metal is added to water.



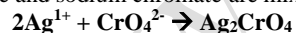
11. Solid calcium carbonate is heated.



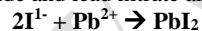
12. A piece of solid bismuth is heated strongly in oxygen.



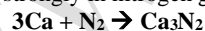
13. Solutions of silver nitrate and sodium chromate are mixed.



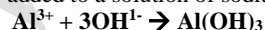
14. Solutions of sodium iodide and lead nitrate are mixed.



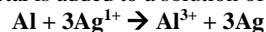
15. Calcium metal is heated strongly in nitrogen gas.



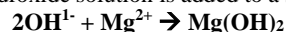
16. Solid aluminum oxide is added to a solution of sodium hydroxide.



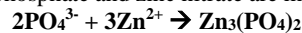
17. A piece of aluminum metal is added to a solution of silver nitrate.



18. An excess of sodium hydroxide solution is added to a solution of magnesium nitrate.



19. Solutions of potassium phosphate and zinc nitrate are mixed.



20. A mixture of powdered iron(III) oxide and powdered aluminum metal is heated strongly.

