

Chapter 6 Reaction Practice II

For each of the following reactions, write a balanced equation for the reaction. Coefficients should be in terms of lowest whole numbers.

1. Magnesium metal is burned in nitrogen gas.
2. Lead foil is immersed in silver nitrate solution.
3. A solution of ammonium sulfate is added to a solution of barium hydroxide.
4. Ethanol(C_2H_5OH) is completely burned in air.
5. Solutions of zinc sulfate and sodium phosphate are mixed.
6. Solutions of silver nitrate and lithium bromide are mixed.
7. A solution of ammonium thiocyanate is added to a solution of iron(II) chloride.
8. Carbon disulfide vapor is burned in excess oxygen.
9. A solution of sodium hydroxide is added to a solution of ammonium chloride.
10. A strip of magnesium is added to a solution of silver nitrate.

11. Ethyne(C_2H_2) gas is burned in air.
12. Solid calcium is added to warm water.
13. Propanol(C_3H_7OH) is burned completely in air.
14. A piece of lithium metal is dropped into a container of nitrogen gas.
15. A solution of sodium sulfide is added to a solution of zinc nitrate.
16. Chlorine gas is bubbled into a solution of potassium iodide.
17. Sodium metal is added to water.
18. Solid calcium carbonate is heated.
19. A piece of solid bismuth is heated strongly in oxygen.
20. Solutions of silver nitrate and sodium chromate are mixed.
21. Solutions of sodium iodide and lead nitrate are mixed.
22. Calcium metal is heated strongly in nitrogen gas.