

Name _____

Honors Chemistry

___/___/___

Chemistry Vocab

Identify each of the following as exothermic or endothermic.

1. Define endothermic -

2. Define exothermic -

3. _____ A puddle of water **freezing** on a dreary day in December

4. _____ Dew **condensing** on a piece of grass on a early September morning

5. _____ **Evaporation** of the tears off the faces of the Washington Redskins after they got their butts kicked by the Philadelphia Eagles on a Sunday in September.

6. _____ **Sublimation** of dry ice on a stage in an auditorium in Terrell, Texas in June 1961

7. _____ **Melting** of ice in a glass on a table at a wedding reception in Virginia at 8:43 PM

8. _____ A candle **melting** on a desk in a classroom on September 10th at 7:50 AM

Identify each of the following as a physical or chemical reaction.

9. Define physical change -

10. Define chemical change -

11. _____ Evaporation of the dew on a piece of grass on a beautiful October morning

12. _____ Separation of water into hydrogen gas and oxygen gas on the 18th of August

13. _____ Separation of the compound carbon dioxide into carbon and oxygen

14. _____ The tearing up of the Redskins playbook into a thousand pieces after realizing that it is ineffective against a powerhouse football team like the Philadelphia Eagles.

15. _____ Burning of illegal documents in an office on a sunny day in late November.

16. _____ Melting of an ice cube in a plastic cup in an uncomfortably warm room on September 18th.

Identify each of the following as a compound, monoatomic element, or molecular element

17. _____ No

22. _____ C₆₀

18. _____ Fe

23. _____ NaCl

19. _____ N₂

24. _____ SO₃

20. _____ NO

25. _____ C₆H₁₂O₆

21. _____ O₃

26. _____ Pt

Separating Mixtures
Part I – Matching

- | | |
|-------------------------------|--|
| 1. _____ chromatography point | a. separation technique which leaves behind the substance with the highest boiling point |
| 2. _____ distillation | b. separation technique where magnetic substances are removed using a magnet |
| 3. _____ centrifuge | c. separation of the substances in a solution by their density |
| 4. _____ evaporation | d. separation of the substances in a mixture based on their boiling point |
| 5. _____ magnetism | e. separation technique where electricity is used to bring about a chemical change |
| 6. _____ electrolysis | f. separation of the substances in a solution based on color |
| 7. _____ decanting | g. separation of the substances in a mixture using a barrier |
| 8. _____ filtration | h. separation technique where a substance changes directly from a gas to a solid |
| 9. _____ sublimation | i. separation of the substances in a heterogeneous mixture by pouring off the top liquid and leaving the solids behind |
| 10. _____ deposition | j. separation technique where a substance changes directly from a solid to a gas |

For each of the following, determine which separation techniques would be best. You will use some separation techniques more than once.

- _____ separation of salt from salt water
- _____ separation of blood into different parts based on their densities
- _____ separation and identification of the substances in a liquid based on color
- _____ separation of 2 substances in a liquid solution that have different densities
- _____ separation of 2 substances in a liquid with differences in volatility (boiling point)
- _____ separation of water into hydrogen gas and oxygen gas
- _____ separation of water from a 2 phase heterogeneous oil - water mixture
- _____ separation of small particles from larger particles using a barrier
- _____ separation of iodine from an iodine – salt mixture
- _____ separation of iron from an iron – sulfur mixture
- _____ separation of tea leaves from a cup of tea
- _____ separation where a water vapor condenses to solid without becoming a liquid