

Name _____

Honors Chemistry

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Chapter 6 Reaction Practice II

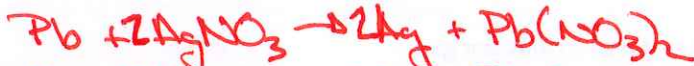
For each of the following reactions, write a balanced equation for the reaction. Coefficients should be in terms of lowest whole numbers.

1. Magnesium metal is burned in nitrogen gas.



Synthesis

2. Lead foil is immersed in silver nitrate solution.



Single



3. A solution of ammonium sulfate is added to a solution of barium hydroxide.



double

4. Ethanol($\text{C}_2\text{H}_5\text{OH}$) is completely burned in air.



combustion



5. Solutions of zinc sulfate and sodium phosphate are mixed.



double

6. Solutions of silver nitrate and lithium bromide are mixed.



double

7. A solution of ammonium thiocyanate is added to a solution of iron(II) chloride.



double

8. Carbon disulfide vapor is burned in excess oxygen.



combustion

9. A solution of sodium hydroxide is added to a solution of ammonium chloride.



double

10. A strip of magnesium is added to a solution of silver nitrate.



Single

11. Ethyne(C_2H_2) gas is burned in air.



Combustion

12. Solid calcium is added to warm water.



Single

13. Propanol(C_3H_7OH) is burned completely in air.



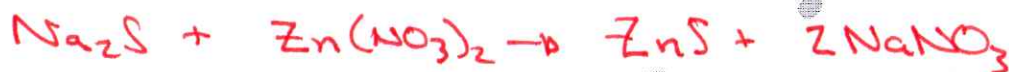
Combustion

14. A piece of lithium metal is dropped into a container of nitrogen gas.



Synthesis

15. A solution of sodium sulfide is added to a solution of zinc nitrate.



double

16. Chlorine gas is bubbled into a solution of potassium iodide.



Single

17. Sodium metal is added to water.



Single

18. Solid calcium carbonate is heated.



decomposition

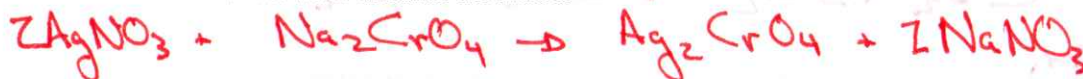
19. A piece of solid bismuth is heated strongly in oxygen.



Synthesis

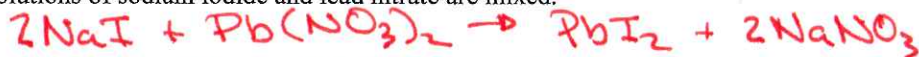


20. Solutions of silver nitrate and sodium chromate are mixed.



double

21. Solutions of sodium iodide and lead nitrate are mixed.



double

22. Calcium metal is heated strongly in nitrogen gas.



Synthesis